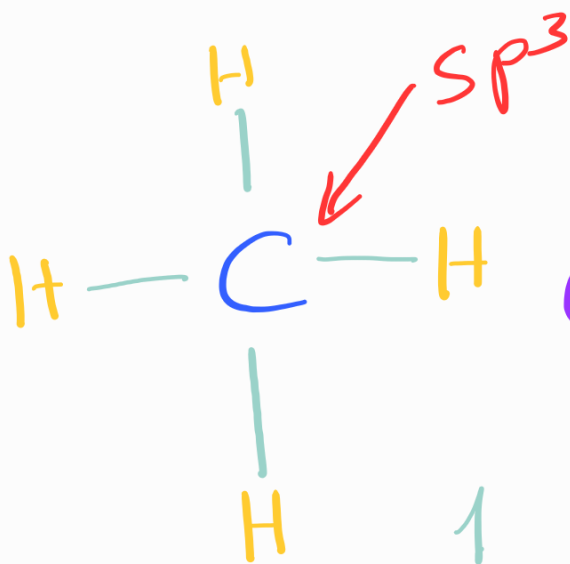
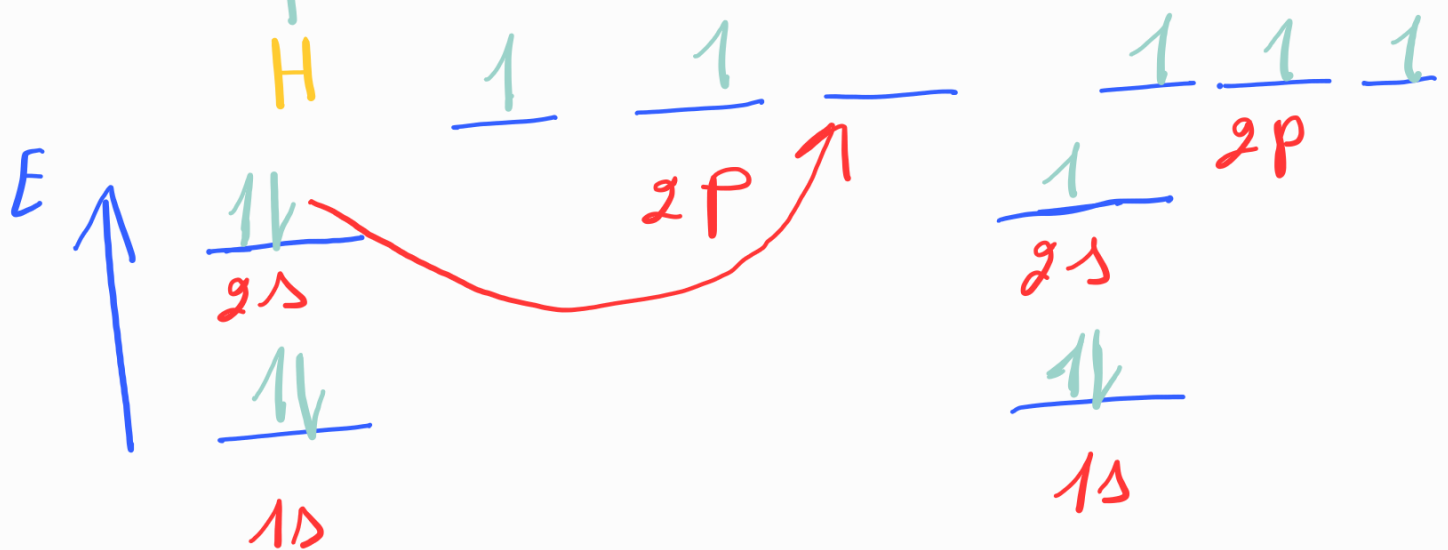


Hybridation



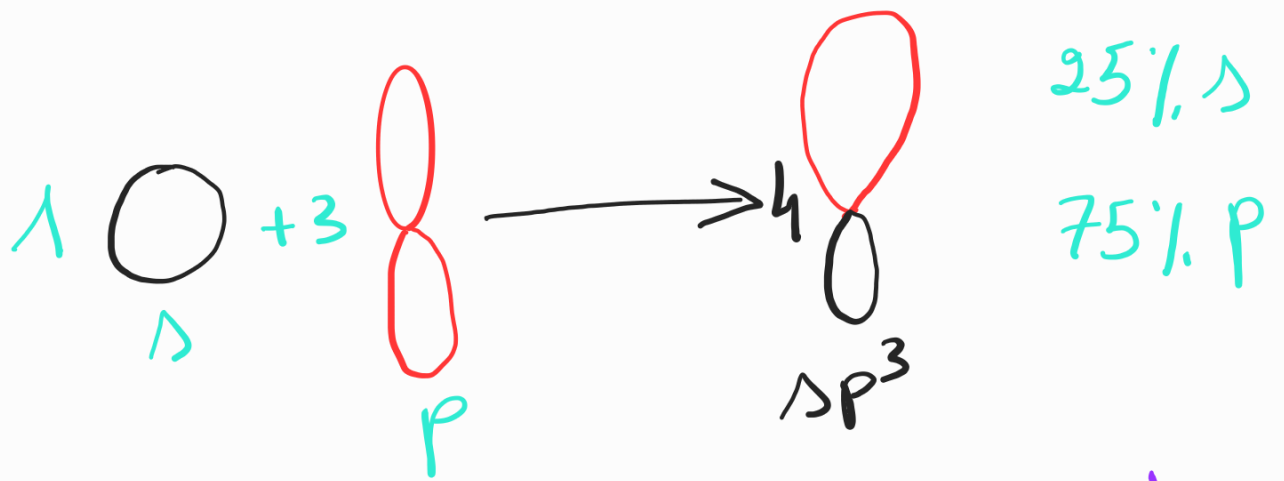
@ physics and chemistry club



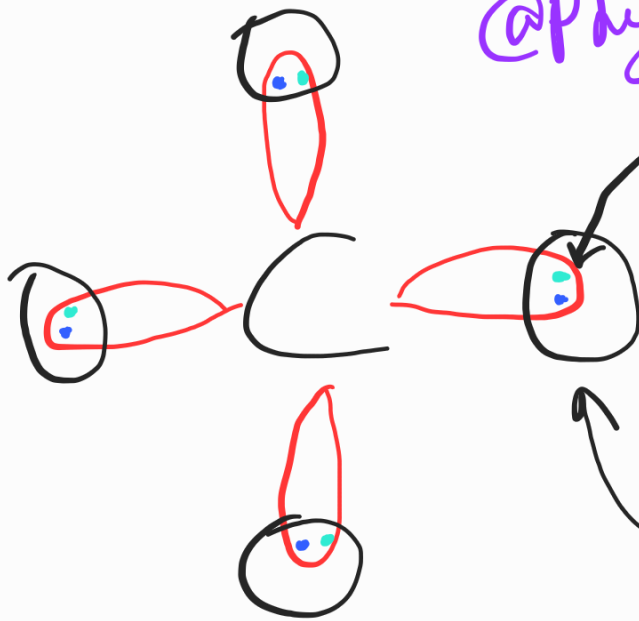
hybridation sp^3 :

$\frac{1}{s}$	$\frac{1}{p}$	$\frac{1}{p}$	$\frac{1}{p}$
s	p	p	p
sp^3	sp^3	sp^3	sp^3

- 4 orbitales sp^3 équivalentes en énergie
- Chacune contient $1e^-$



@Physics and Chemistry Club.



mise en commun des
 e⁻ dans cette tête de
 chevauchement.

chevauchement entre
 orbitales hybrides sp^3
 du carbone et orbitales
 $1s$ de l'hydrogène.